

Acid 1/2 reactions

ex



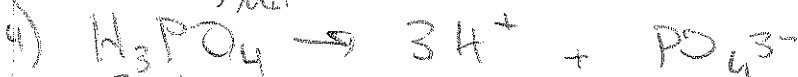
$$50 \text{ gr} / 128 \text{ g/mol} = 0.39 \text{ mol}$$



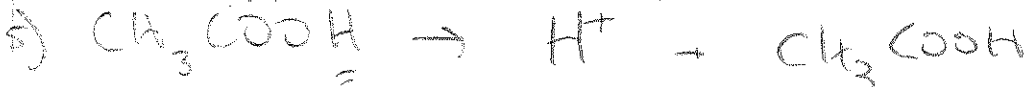
$$50 \text{ gr} / 63 \text{ g/mol} = 0.79 \text{ mol}$$



$$50 / 36.5 \text{ gr/mol} = 1.37 \text{ mol}$$



$$50 \text{ gr} / 98 \text{ gr/mol} = 0.51 \text{ mol} \times 3 = 1.53 \text{ mol}$$



$$50 \text{ gr} / 60 \text{ g/mol} = 0.83 \text{ mol}$$

7) 1.95

2) 3.95

3) 6.85

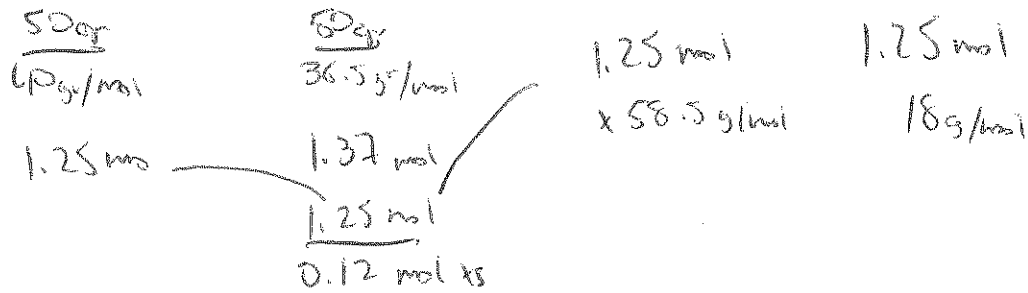
4) 7.65

5) 4.15

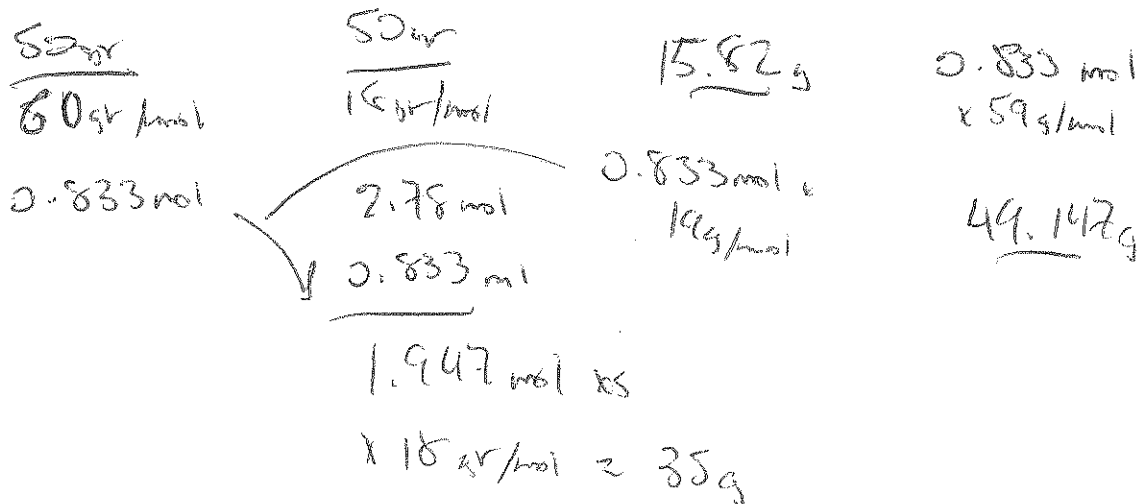
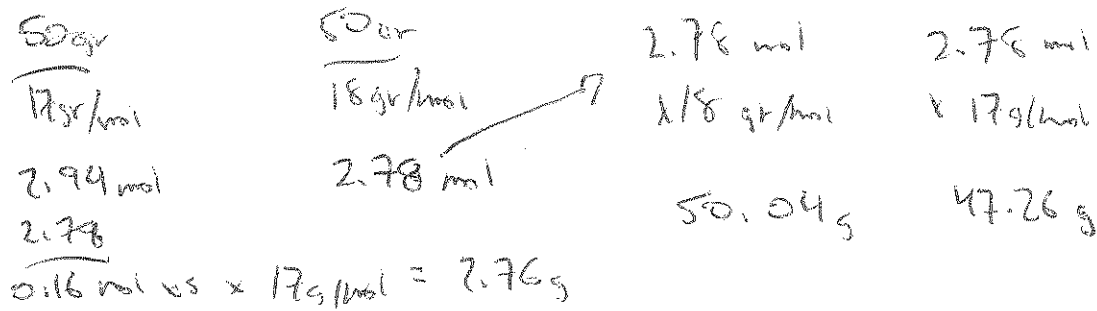
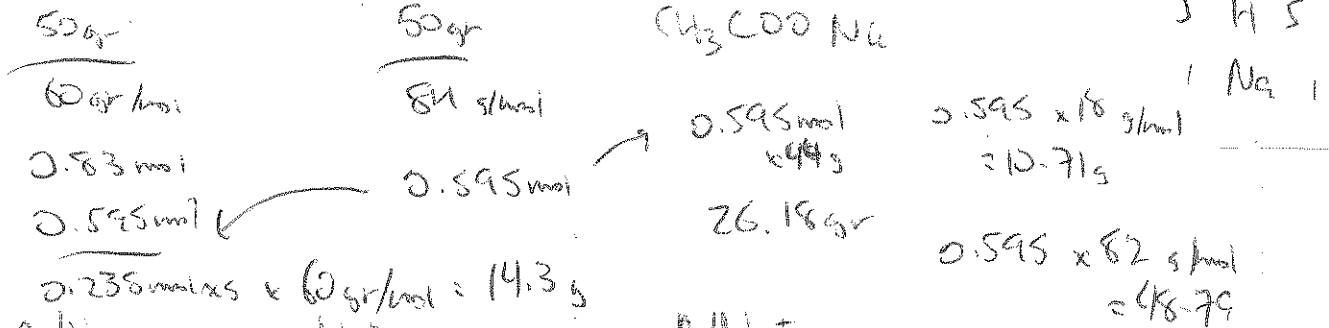
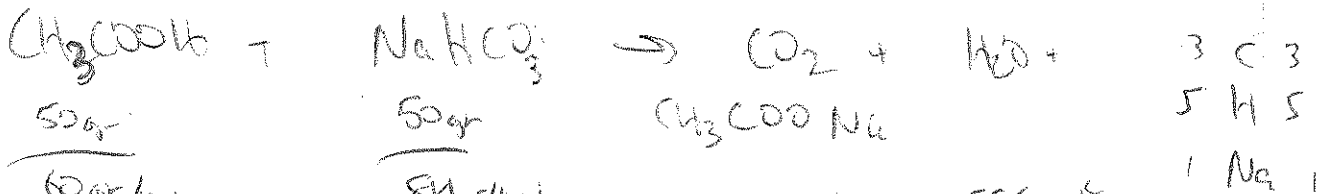
6) given 50 gr of each of the above, how much H^+ can be liberated

7) if the amount of H^+ from above is dissolved in 200 ml of H_2O what is the concentration of H^+

Given 50 gr of each, calc products and
xs reagent



✓ thus



Balance + solve for limiting + excess reagents given 75 gr of each reactant

